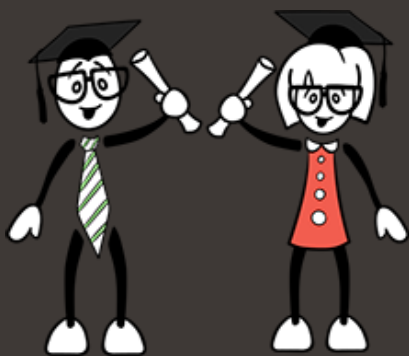


**ICSE CLASS 10**

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**Std. X ICSE**

**Chemistry**

**Marks: 80**

Reading time: 10 Min. Marks will be deducted for untidy Presentation.

**Section I**

**Marks: 40**

1 a) From the list given below select the words require to correctly complete blanks. [5]  
[Ammonia, ammonium, carbonate, carbon dioxide, hydrogen, hydronium, hydroxide, precipitate, salt, water]

1) A solution M turns blue litmus red, so it must contain i \_\_\_\_\_ ions, another solution O turns red litmus blue and hence, must contain ii \_\_\_\_\_ ions.


2) When solution M and O are mixed together, the products will be iii \_\_\_\_\_ and iv \_\_\_\_\_.

3) If a piece of magnesium was put into a solution M, (v) \_\_\_\_\_ gas would be evolved.

b) Identify the gas evolved in the following reactions when:

- i. Sodium propionate is heated with soda lime.
- ii. Potassium sulphite is treated with dilute hydrochloric acid.
- iii. Sulphur is treated with concentrated nitric acid.
- iv. A few crystals of  $\text{KNO}_3$  are heated in a hard glass test tube.
- v. Concentrated hydrochloric acid is made to react with manganese dioxide.

c) Give one word or phrase for the following. [5]

- i. The ratio of the mass of a certain volume of gas to the mass of an equal volume of hydrogen under the same conditions of temperature and pressure.
- ii. Formation of irons from molecules.
- iii. Electrolytic deposition of a superior metal on a baser metal.
- iv. \_\_\_\_\_
- v. Hydrocarbon containing a  functional group.
- vi. The amount of energy released when on atom in the gaseous state accepts an electrons to form an anion.

d) Match the options A to E with the statements (i) to (v). [5]

- A) Alkynes  
B) alkane  
C) Iron  
D)  $6.023 \times 10^{23}$   
E) metal

- i) No of molecules in  $22.4 \text{ dm}^3$  of  $\text{CO}_2$  at s.t.q.  
ii) An element with electronic configuration 2,8,8,3  
iii)  $\text{C}_n\text{H}_{2n+2}$   
iv)  $\text{C}_n\text{H}_{2n-2}$

- v) The metal that forms two types of irons
- e) Write balanced equations for the following. [5]
- 1) Action of heat on a mixture of copper and concentrated nitric acid.
  - 2) Action of warm water on magnesium nitride.
  - 3) Action of concentrated sulphuric acid carbonic.
  - 4) Action of dilute hydrochloric acid on sodium sulphide.
  - 5) Preparation of ethane from sodium propionate.
- f) Distinguish between the following pairs of compounds using the test given within brackets. [5]
- i) Iron (II) sulphate and iron (III) sulphate (using ammonium hydroxide)
  - ii) A lead salt and a zinc salt (using excess ammonium hydroxide)
  - iii) Sodium nitrate and Sodium sulphite (using dil sulphuric acid)
  - iv) Dilute sulphuric acid and dilute hydro chloric acid (using barium chloride solution)
  - v) Ethane and ethene (using alkaline potassium permanganate solution)
- g) [5]
- i) Oxygen oxidizes ethyne to  $\text{CO}_2$  & water as shown by the equations.
- ii)  $2\text{C}_2\text{H}_2 + \text{SO}_2 \rightarrow 4\text{CO}_2 + 2\text{H}_2\text{O}$

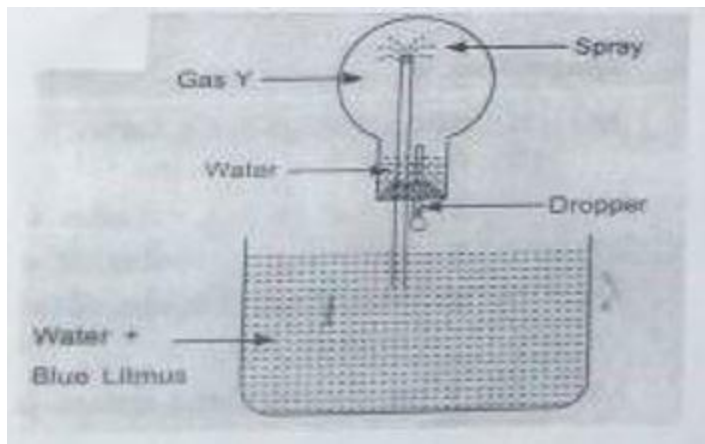
What volume of ethyne gas at st.p is required to produce  $84\text{ dm}^3$  of carbon dioxide at st.p?

- ii) A compound made up of two elements X and Y has an empirical formula  $\text{X}_2\text{Y}$ . If the atomic weight of X is 10 and that Y is 5 and the compound has a vapour density 25, find its molecular formula.

## Section II [40 M]

Attempt any four

- a) State your observation in each of the following cases.
- b)
- i) When dil HCL acid is added to sodium carbonate crystals.
  - ii) When excess sodium hydroxide is added to calcium nitrate solution.
  - iii) All the cathode when acidified aqueous copper sulphate solution is electrolyzed with copper electrodes.
  - iv) When moist starch iodide paper is introduced into chlorine gas.
- c) State the figure below and answer the questions the follow:



- i) Identify the gas Y
- ii) What property of gas y does this experienced demonstrate.
- iii) Name another gas which has the same property and can be demonstrated through this experiment

**Q2**

- i. Name the other ion formed when ammonia dissolve in water.
- ii. Give one test that can be used to detect.

**Q 3) a) State the conditions required for the following reactions to take place. [5]**

- 1) Catalytic hydrogenation of ethyne.
- 2) Preparation of ethyne from ethylene dibromide.
- 3) Catalytic oxidation of ammonia to ammonia to nitric oxide.
- 4) Any two conditions for the conversion of Sulphur dioxide to Sulphur trioxide.

**b) State the main components of the following alloys: - [3]**

- 1) Brass
- 2) Duralumin
- 3) Bronze

**c) Give balanced equations for the following: - [2]**

- 1) Laboratory preparation of nitric acid
- 2) Preparation of ethanol from nmochlorethane and aq sodium hydroxide.

**Q 4) a) Give the structural formula of the following: - [4]**

- i. Ethanol
- ii. 1-proparal
- iii. Ethannic acid acid
- iv. 1,2, dichloroethane

b) Draw the structure of the stable positive ion formed when an acid dissolves in water. [2]

c) State the inference drawn from the following observations:- [4]

i. On carrying out the flame test with a salt P a brick red flame was obtained what is the cation in P?

ii. A gas Q turns moist lead acetate paper silvery black. Identify the gas Q.

iii. A gas PH of liquid R is 10. What kind of substance is R?

iv. Salt is prepared by reacting dilute sulphuric acid with copper oxide. Identify S

Q 5) a) Name the following: [3]

1) The property possessed by metals by which they can be beaten in to sheets.

2) A compound added to lower the fusion temperature of electrolytic bath in the extraction of aluminium.

3) The ore of zinc containing its sulphide.

b) Give one equation each to show the following properties of sulphuric acid:- [3]

1) Dehydrating property

2) Acid nature

3) As a non volatile acid.

c) Give balanced chemical equations to prepare the following salts: [4]

1) Lead sulphate from lead carbonate

2) Sodium sulphate using dilute sulphuric acid

3) Copper chloride using copper carbonate.

Q 6) a) 1 – State Avogadro's law [4]

2) – A cylinder contains 68 g of ammonia gas at S.t.P.

a) What is the volume occupied by the gas?

b) How many moles of ammonia are present in the cylinder?

c) How many molecules of ammonia are present in the cylinder? [N = 14 H = 1]

b) 1) Why do covalent compounds exist as gases, liquids or soft solids? [3]

2) Which electrode: anode or cathode is the oxidizing electrode? Why?

c) Name the kind of particles present in: [3]

1) Sodium Hydroxide Sol<sup>n</sup>

2) Carbonic acid

3) Sugar solution.

Q 7 a) An element Z has atomic number 16. [5]

Answer the following questions on Z.

- i. State the period and group to which Z belongs.
- ii. Is Z a metal or a non metal?
- iii. State the formula between Z and Hydrogen
- iv. What kind of a compound is this.

b) M is a metal above hydrogen in the activity series and its oxide has the formula  $M_2O$ . This oxide when dissolved in water forms the corresponding hydroxide which is a good conductor of electricity. [5]

In the above context answer the following.

- 1) What kind of combination exists between M and O?
- 2) How many electrons are there in the outer most shell of M?
- 3) Name the group to which M belongs?
- 4) State the reaction taking place at the cathode
- 5) Name the product at the anode.

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